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AP Chemistry Chapter 12 Answers Zumdahl12.49For elementary reactions, the rate law can be written using the coefficients in the balanced equation to determine orders.a. Rate = $k[\text{CH}_3\text{NC}]$ b. Rate = $k[\text{O}_3][\text{NO}]$ c. Rate = $k[\text{O}_3]$ d. Rate = $k[\text{O}_3][\text{O}]$ 12.50The observed rate law for this reaction is: Rate = $k[\text{NO}]^2[\text{H}_2]$. For a mechanism to be plausible, the sum of all the steps must give the overall balanced equation...

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AP Chemistry Chapter 14 Answers Zumdahl14.27a. $\text{HClO}_4(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{H}_3\text{O}^+(\text{aq}) + \text{ClO}_4^-(\text{aq})$. Only the forward reaction is indicated since HClO_4 is a strong acid and is basically 100% dissociated in water. For acids, the dissociation reaction is commonly written without water as a reactant. The common abbreviation for this reaction is: $\text{HClO}_4(\text{aq}) \rightleftharpoons \text{H}^+(\text{aq}) + \text{ClO}_4^-(\text{aq})$. This reaction is also called the K_a re...

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AP Chemistry Chapter 5 Answers Zumdahl5.27a. $4.8 \text{ atm} \times 760 \text{ mmHg} / 1 \text{ atm} = 3.6 \times 10^3 \text{ mmHg}$. $3.6 \times 10^3 \text{ mmHg} \times 1 \text{ torr} / 1 \text{ mmHg} = 3.6 \times 10^3 \text{ torr}$. $4.8 \text{ atm} \times 1.013 \times 10^5 \text{ Pa} / 1 \text{ atm} = 4.9 \times 10^5 \text{ Pa}$. $4.8 \text{ atm} \times 14.7 \text{ psi} / 1 \text{ atm} = 71 \text{ psi}$ 5.296.5 cm x 10 mm / 1 cm = 65 mmHg = 65 torr; 65 torr x 1 atm / 760 torr = $8.6 \times 10^{-2} \text{ atm}$ 6.6 x 10⁻² atm x 1.013 x 10⁵ Pa / 1 atm = $8.7 \times 10^4 \text{ Pa}$ 5.31If the levels of Hg in each arm of the m...

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