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AP Chemistry Chapter 5 Answers Zumdahl 5.27a.  $4.8 \text{ atm} \times 760 \text{ mmHg} / 1 \text{ atm} = 3.6 \times 10^3 \text{ mmHg}$ .  $3.6 \times 10^3 \text{ mmHg} \times 1 \text{ torr} / 1 \text{ mmHg} = 3.6 \times 10^3 \text{ torr}$ .  $4.8 \text{ atm} \times 1.013 \times 10^5 \text{ Pa} / 1 \text{ atm} = 4.9 \times 10^5 \text{ Pa}$ .  $4.8 \text{ atm} \times 14.7 \text{ psi} / 1 \text{ atm} = 71 \text{ psi}$  5.29  $6.5 \text{ cm} \times 10 \text{ mm} / 1 \text{ cm} = 65 \text{ mmHg} = 65 \text{ torr}$ ;  $65 \text{ torr} \times 1 \text{ atm} / 760 \text{ torr} = 8.6 \times 10^{-2} \text{ atm}$  6.6  $6 \times 10^{-2} \text{ atm} \times 1.013 \times 10^5 \text{ Pa} / 1 \text{ atm} = 8.7 \times 10^4 \text{ Pa}$  5.31 If the levels of Hg in each arm of the m...

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AP Chemistry Chapter 9 Answers Zumdahl 9.15  $\text{H}_2\text{O}$  has  $2(1) + 6 = 8$  valence electrons  $\text{H}_2\text{O}$  has a tetrahedral arrangement of the electron pairs about the O atom that requires  $\text{sp}^3$  hybridization. Two of the four  $\text{sp}^3$  hybrid orbitals are used to form bonds to the two hydrogen atoms and the other two  $\text{sp}^3$  hybrid orbitals hold the two lone pairs of oxygen. The two O-H bonds are formed from overlap of the  $\text{sp}^3$  hybrid...

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